Ionic conductivity and electric relaxation in the poly(NN-dimethylacrylamide)-LiClO₄ system

B.-E. Mellander
Department of Physics, Chalmers University of Technology, Göteborg, Sweden

J. R. Stevens
Department of Physics, University of Guelph, Guelph, Ontario N1G 2W1, Canada

Z. Naturforsch. 43a, 1044–1046 (1988); received September 8, 1988

The poly(NN-dimethylacrylamide)-LiClO₄ system has been investigated regarding ionic conductivity and electric relaxation. This material may contain up to 29 mol% LiClO₄ while still being an amorphous and transparent ionic conductor. The conductivity for this composition is of the order of $10^{-9}$ (ohm cm)$^{-1}$ at room temperature while it increases to $7\times10^{-7}$ (ohm cm)$^{-1}$ at 100°C. The activation energy according to the conductivity measurements is 1.1 eV. The system poly(methylmethacrylate)-LiCF₃SO₃ was also investigated but only very low ionic conductivities were detected.

Key words: Ionic conductivity, electric relaxation dimethylacrylamide, lithium perchlorate, polymer electrolyte.

Introduction

During the last few years the interest for ion conducting polymers has grown rapidly [1–5]. The systems that have been studied most have been based on poly(ethylene oxide) (PEO) and it has been shown that a high ionic conductivity occurs mainly above the glass transition temperature in elastomeric materials, while crystalline materials have low ionic conductivities [3].

In many ion conducting glasses the ionic conductivity increases with increasing amounts of the network modifier [6] i.e. a high conductivity may be obtained by adding a rather large amount of an inorganic salt to the glass. In the present paper we report on studies of the ionic conductivity when LiClO₄ is added to a transparent amorphous polymer, poly(NN-dimethylacrylamide) (PDMA). The glass transition temperature for dominantly atactic PDMA is 120°C [7]. The DMA monomer unit, see Fig. 1, contains one oxygen in the side group rather than in the polymer backbone as is the case for poly(ethylene oxide). The ionic conductivity of the system poly(methyl methacrylate) (PMMA)-LiCF₃SO₃ was also studied.

The electrical relaxation has been studied using electric modulus analysis [8–13]. The complex electric modulus $M$ is equal to the inverse of the complex permittivity,

$$M = (\varepsilon')^{-1} = M' + jM''$$

and the complex modulus is related to the complex impedance $Z$ through

$$M = j\omega C_0 Z$$

where $C_0$ is the vacuum capacitance of the cell. Electric modulus measurements have earlier been performed mainly for β-alumina and ion conducting glasses, where it has been shown to be a useful method to study materials with mobile ions [8–13].

Experimental

NN-dimethylacrylamide monomer (Fluka, 98% pure) was freeze-dried and distilled under vacuum. Dried LiClO₄ (Ventron, 99.5% pure) was dissolved in the liquid monomer at 90°C, and the samples was then freeze-dried again to avoid traces of water. The polymerization was performed at a temperature of

$\text{H} -\text{C} = \text{O}$

$\text{H} \quad \text{H}$

$\text{N}-\text{CH}_3$

$\text{CH}_3$

Fig. 1. The poly(NN-dimethylacrylamide) repeat unit.

Reprint requests to Dr. B.-E. Mellander, Department of Physics, Chalmers University of Technology, S-41296 Göteborg, Sweden.
40 °C during a few weeks. The PMMA-LiCF₃SO₃ samples were prepared in a similar way using 97% pure LiCF₃SO₃ (Aldrich Chemical Company Inc., England).

The ionic conductivity measurements were performed using complex impedance analysis [14]. The complex impedance and the electrical modulus were measured in the frequency range 100 Hz to 100 kHz using a computer controlled HP4274A LCR meter. The samples were about 7 mm in diameter and 1–2 mm thick. Graphite electrodes were painted on the sample surfaces using graphite dissolved in alcohol (DAG-580, Acheson, Holland). Another type of electrode that was also used for the PDMA samples was a thin platinum foil which was made to stick to the sample surfaces if these were slightly moistened using a small drop of ethanol. After the application of the graphite or platinum electrodes the sample was dried in vacuum before the measurement was started, since even small amounts of residual solvent may influence the results. Good agreement was obtained between results using the two kinds of electrodes. Wide angle X-ray diffraction measurements have been done using a Philips X-ray diffractometer employing Ni filtered CuKα radiation. All measurements were performed in a nitrogen atmosphere or in vacuum.

Results and Discussion

The electrical conductivity of PDMA samples with 9.4 and 13.5 mol% LiClO₄ was negligible, less than $10^{-9}$ (ohm cm)$^{-1}$ for temperatures between 25 and 115 °C. For samples with 29.4 mol% LiClO₄, however, the conductivity was of the order of $10^{-9}$ (ohm cm)$^{-1}$ at ambient temperature and increased to $7 \times 10^{-7}$ (ohm cm)$^{-1}$ at 100 °C. We have not, so far, been able to produce homogenous non-crystalline samples with higher LiClO₄ concentration. For all other samples, however, wide angle X-ray diffraction measurements confirmed that the materials were completely amorphous. The temperature dependence of the electrical conductivity for a sample with 29.4 mol% LiClO₄ is shown in Figure 2. The average activation energy for this composition, determined from several runs, is 1.07 ± 0.04 eV. The electronic contribution to the total conductivity has been investigated by a dc-polarization method. No detectable electronic conductivity was, however, found.
The real and imaginary parts of the electric modulus have been studied as a function of frequency and temperature. In Fig. 3 the temperature dependence of the imaginary part of the complex modulus is shown. The relaxation time calculated from the relaxation peak is shown as a function of temperature in Fig. 4, where results obtained from plots of the imaginary part of the modulus as a function of log (frequency) as well as of temperature are included. The average activation energy calculated from the relaxation time plots is 1.2 eV. The difference between the activation energies obtained from conductivity and relaxation measurements is, however, not significant since the error in the relaxation measurements is relatively large. The width at half-height of the relaxation peaks as a function of frequency exceeds the Debye value of 1.144 decades as is the case for many other ionic conductors. Typical values for the width are of the order of 2.4 decades, and the peak is unsymmetrical, broader on the high frequency side than on the low frequency side.

The PMMA-LiCF₃SO₃ samples were all transparent and amorphous, but the ionic conductivities were very low, below 10⁻⁸ even at temperatures as high as 150 °C. The compositions investigated covered the range 9 to 29 mol% LiCF₃SO₃.

Acknowledgements

We are indebted to I. Albinsson and P. A. Svanteson for performing the X-ray analysis of our samples. Financial support from the Swedish Board of Technical Development, the Natural Sciences and Engineering Research Council of Canada, Carl Tryggers Stiftelse för vetenskaplig forskning and Stiftelsen Wilhelm och Martina Lundgrens Vetenskapsfond is gratefully acknowledged.